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**KOLEJ YAYASAN PELAJARAN JOHOR  
PEPERIKSAAN AKHIR**

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NAMA KURSUS : TERMODINAMIK  
KOD KURSUS : DKM 3203  
PEPERIKSAAN : OKTOBER 2017  
MASA : 3 JAM

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**ARAHAN KEPADA CALON**

1. Kertas soalan ini mengandungi **SATU (1)** bahagian. Bahagian A (100 markah)
2. Calon tidak dibenarkan membawa masuk sebarang peralatan ke dalam bilik peperiksaan kecuali dengan kebenaran pengawas peperiksaan.
3. Sila pastikan bahan-bahan berikut diperoleh untuk sesi peperiksaan ini:
  - i. Kertas Soalan
  - ii. Buku Jawapan
  - iii. Jadual Stim

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**JANGAN BUKA KERTAS SOALANINI SEHINGGA DIBERITAHU**

**KERTAS SOALANINI MENGANDUNG~~6~~ HALAMAN BERCETAK TERMASUK MUKA HADAPAN**

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Bahagian ini mengandungi LIMA (5) soalan.

Jawab EMPAT daripada LIMA soalan. Jawab dalam Buku Jawapan.

### SOALAN 1

#### QUESTION 1

- a. Tukarkan unit berikut:

*Convert the following units:*

- (i)  $25 \text{ g/mm}^3$  kepada  $\text{kg/m}^3$   
 $25 \text{ g/mm}^3$  to  $\text{kg/m}^3$
- (ii)  $6\text{N/cm}^2$  kepada  $\text{kN/m}^2$   
 $6\text{N/cm}^2$  to  $\text{kN/m}^2$
- (iii)  $380 \text{ km/j}$  kepada  $\text{cm/minit}$   
 $380 \text{ km/hr}$  to  $\text{cm/minute}$

(9 markah)  
(9 marks)

- b. Dengan bantuan gambar rajah, jelaskan maksud tekanan mutlak, tekanan atmosfera, dan tekanan tolok.

*With the aid of diagram, explain the absolute pressure, atmospheric pressure, and gauge pressure.*

(6 markah)  
(6 marks)

- c. Stim pada tekanan  $1000 \text{ kPa}$  mempunyai tenaga dalamnya  $2480 \text{ kJ/kg}$ . Kirakan:

- (i) pecahan kekeringan  
(ii) isi padu tentu  
(iii) entalpi tentu

Kemudian, lakar dan tandakan titik pecahan kekeringan pada rajah P-v.

*Steam at  $1000 \text{ kPa}$  as the specific internal energy  $2480 \text{ kJ/kg}$ . Find:*

- (i) dryness fraction  
(ii) specific volume  
(iii) specific enthalpy

*Then, sketch and locate the dryness fraction on the P-v diagram.*

(10 markah)  
(10 marks)

## SOALAN 2

## QUESTION 2

- a. Tentukan isi padu yang dipenuhi oleh 1 kg stim yang berada pada tekanan  $0.85 \text{ MN/m}^2$  dengan pecahan kekeringan 0.97. Stim tersebut dikembangkan secara adiabatik ke tekanan  $0.175 \text{ MN/m}^2$  berdasarkan hukum  $PV^{1.13}=\text{pemalar}$ . Tentukan pecahan kekeringan stim akhir dan perubahan tenaga dalam stim semasa pengembangan.

*Determine the volume occupied by 1 kg of steam at a pressure  $0.85 \text{ MN/m}^2$  with a dryness fraction of 0.97. This volume is expanded adiabatically to a pressure of  $0.175 \text{ MN/m}^2$  with the law of expansion  $PV^{1.13}=\text{constant}$ . Determine the final dryness fraction of the steam and the change of internal energy of the steam during the expansion.*

(15 markah)

(15marks)

- b. Berbandukan jadual stim, pada tekanan  $2.75 \text{ MN/m}^2$ , tentukan:

- (i) suhu tepu
- (ii) entalpi cair tentu
- (iii) entalpi tentu penyejatan
- (iv) entalpi tentu stim tepu kering

*According to steam table, at pressure of  $2.75 \text{ MN/m}^2$  determine:*

- (i) saturation temperature
- (ii) specific liquid enthalpy
- (iii) specific enthalpy of evaporation
- (iv) specific enthalpy of dry saturated steam

(10 markah)

(10 marks)

**SOALAN 3****QUESTION 3**

- a. Lakarkan gambar rajah suhu-entalpi dan labelkan kawasan entalpi tepu, entalpi panas lampau, dan kawasan cecair tepu.

*Sketch the temperature-enthalpy diagram and label the area of saturated enthalpy, superheated enthalpy, and saturated steam.*

(5 markah)

(5 marks)

- b. Sekumpulan muncung tumpu-capah dibekalkan stim pada tekanan  $2 \text{ MN/m}^2$  dan suhu  $325^\circ\text{C}$ . Pengembangan tepu lampau menurut hukum  $PV^{1.3}=\text{pemalar}$  berlaku di dalam muncung ke tekanan keluar  $0.36 \text{ MN/m}^2$ . Stim dibekalkan pada kadar  $7.5 \text{ kg/s}$ . Tentukan:

- (i) luas kerongkong  
(ii) tahap penyejukan di bahagian keluar

*A group of convergent-divergent nozzles are supplied with steam at a pressure of  $2 \text{ MN/m}^2$  and temperature  $325^\circ\text{C}$ . Superstaturated expansion according to the law  $PV^{1.3}=\text{constant}$  occurs in the nozzle down to an exit pressure of  $0.36 \text{ MN/m}^2$ . Steam is supplied at the rate of  $7.5 \text{ kg/s}$ . Determine:*

- (i) throat area  
(ii) the degree of undercooling at exit

(20 markah)

(20 marks)

**SOALAN 4****QUESTION 4**

- a. Terangkan Hukum Kedua Termodinamik.

*Explain about Second Law of Thermodynamics.*

(3 markah)

(3 marks)

- b. Stim berjisim 4.5 kg berada pada tekanan awal  $3 \text{ MN/m}^2$  dan suhu awal  $300^\circ\text{C}$ . Stim tersebut mengembang secara boleh balik kepada tekanan yang baru iaitu  $0.1 \text{ MN/m}^2$  dengan faktor kekeringan 0.96. Proses pengembangan stim dilihat sebagai satu garisan lurus apabila diplotkan di atas carta suhu-entropi.

Tentukan:

- (i) haba terpindah semasa proses pengembangan
- (ii) kerja terlaku semasa proses pengembangan

*4.5 kg of steam has an initial pressure of  $3 \text{ MN/m}^2$  and temperature of  $300^\circ\text{C}$ . The steam then expands reversibly to a new pressure of  $0.1 \text{ MN/m}^2$  with a dryness factor of 0.96. The expansion of the steam appears as a straight line when plotted on a temperature-entropy chart. Determine:*

- (i) the heat transfer during the expansion
- (ii) the work done during the expansion

(22 markah)

(22 marks)

## SOALAN 5

### QUESTION 5

- a. Huraikan tentang Prinsip Carnot.

*Elaborate the Carnot Principles.*

(5 markah)

(5 marks)

- b. Satu kilogram udara digunakan dalam Kitar Carnot. Tekanan awal  $1.73 \text{ MN/m}^2$  dan suhu awal  $250^\circ\text{C}$ . Dari keadaan awal, udara dikembangkan secara isoterma menjadi tiga kali isi padu asal dan diikuti pengembangan adiabatik sebanyak enam kali isi padu asal. Mampatan isoterma dan diikuti pula dengan mampatan adiabatik, melengkapkan kitar tersebut. Tentukan: tekanan, isi padu dan suhu di setiap penjuru kitar. Tuliskan jawapan akhir dalam bentuk jadual.

Diberi  $R = 0.29 \text{ kJ/kg.K}$  dan  $\gamma = 1.4$

One kilogram of air is taken through a Carnot Cycle. The initial pressure and temperature of the air are  $1.73 \text{ MN/m}^2$  and  $250^\circ\text{C}$ , respectively. From the initial condition, the air is expanded isothermally to three times its initial volume and then further expanded adiabatically to six times its initial volume. Isothermal compression, followed by adiabatic compression, completes the cycle. Determine the pressure, volume and temperature at each corner of the cycle. Summarize the final answer in table.

Given  $R = 0.29 \text{ kJ/kg.K}$  and  $\gamma = 1.4$

(20 markah)

(20 marks)

[100 MARKAH]

KERTAS SOALAN TAMAT



TABLE A-4

Saturated water—Temperature table (Concluded)

Temp., °C	Sat. press., P <sub>sat</sub> , kPa	Specific volume, m <sup>3</sup> /kg		Internal energy, kJ/kg			Enthalpy, kJ/kg			Entropy, kJ/kg·K		
		Sat. liquid, v <sub>f</sub>	Sat. vapor, v <sub>g</sub>	Sat. liquid, u <sub>f</sub>	Evap., u <sub>fg</sub>	Sat. vapor, u <sub>g</sub>	Sat. liquid, h <sub>f</sub>	Evap., h <sub>fg</sub>	Sat. vapor, h <sub>g</sub>	Sat. liquid, s <sub>f</sub>	Evap., s <sub>fg</sub>	Sat. vapor, s <sub>g</sub>
205	1724.3	0.001164	0.11508	872.86	1723.5	2596.4	874.87	1920.0	2794.8	2.3776	4.0154	6.3930
210	1907.7	0.001173	0.10429	895.38	1702.9	2598.3	897.61	1899.7	2797.3	2.4245	3.9318	6.3563
215	2105.9	0.001181	0.094680	918.02	1681.9	2599.9	920.50	1878.8	2799.3	2.4712	3.8489	6.3200
220	2319.6	0.001190	0.086094	940.79	1660.5	2601.3	943.55	1857.4	2801.0	2.5176	3.7664	6.2840
225	2549.7	0.001199	0.078405	963.70	1638.6	2602.3	966.76	1835.4	2802.2	2.5639	3.6844	6.2483
230	2797.1	0.001209	0.071505	986.76	1616.1	2602.9	990.14	1812.8	2802.9	2.6100	3.6028	6.2128
235	3062.6	0.001219	0.065300	1010.0	1593.2	2603.2	1013.7	1789.5	2803.2	2.6560	3.5216	6.1775
240	3347.0	0.001229	0.059707	1033.4	1569.8	2603.1	1037.5	1785.5	2803.0	2.7018	3.4405	6.1424
245	3651.2	0.001240	0.054656	1056.9	1545.7	2602.7	1061.5	1740.8	2802.2	2.7476	3.3596	6.1072
250	3976.2	0.001252	0.050085	1080.7	1521.1	2601.8	1085.7	1715.3	2801.0	2.7933	3.2788	6.0721
255	4322.9	0.001263	0.045941	1104.7	1495.8	2600.5	1110.1	1689.0	2799.1	2.8390	3.1979	6.0369
260	4692.3	0.001276	0.042175	1128.8	1469.9	2598.7	1134.8	1661.8	2796.6	2.8847	3.1169	6.0017
265	5085.3	0.001289	0.038748	1153.3	1443.2	2596.5	1159.8	1633.7	2793.5	2.9304	3.0358	5.9662
270	5503.0	0.001303	0.035622	1177.9	1415.7	2593.7	1185.1	1604.6	2789.7	2.9762	2.9542	5.9305
275	5946.4	0.001317	0.032767	1202.9	1387.4	2590.3	1210.7	1574.5	2785.2	3.0221	2.8723	5.8944
280	6416.6	0.001333	0.030153	1228.2	1358.2	2586.4	1236.7	1543.2	2779.9	3.0681	2.7898	5.8579
285	6914.6	0.001349	0.027756	1253.7	1328.1	2581.8	1263.1	1510.7	2773.7	3.1144	2.7066	5.8210
290	7441.8	0.001366	0.025554	1279.7	1296.9	2576.5	1289.8	1476.9	2766.7	3.1608	2.6225	5.7834
295	7999.0	0.001384	0.023528	1306.0	1264.5	2570.5	1317.1	1441.6	2758.7	3.2076	2.5374	5.7450
300	8587.9	0.001404	0.021659	1332.7	1230.9	2563.6	1344.8	1404.8	2749.6	3.2548	2.4511	5.7059
305	9209.4	0.001425	0.019932	1360.0	1195.9	2555.8	1373.1	1366.3	2739.4	3.3024	2.3633	5.6657
310	9865.0	0.001447	0.018333	1387.7	1159.3	2547.1	1402.0	1325.9	2727.9	3.3506	2.2737	5.6243
315	10,556	0.001472	0.016849	1416.1	1121.1	2537.2	1431.6	1283.4	2715.0	3.3994	2.1821	5.5816
320	11,284	0.001499	0.015470	1445.1	1080.9	2526.0	1462.0	1238.5	2700.6	3.4491	2.0881	5.5372
325	12,051	0.001528	0.014183	1475.0	1038.5	2513.4	1493.4	1191.0	2684.3	3.4998	1.9911	5.4908
330	12,858	0.001560	0.012979	1505.7	993.5	2499.2	1525.8	1140.3	2666.0	3.5516	1.8906	5.4422
335	13,707	0.001597	0.011848	1537.5	945.5	2483.0	1559.4	1086.0	2645.4	3.6050	1.7857	5.3907
340	14,601	0.001638	0.010783	1570.7	893.8	2464.5	1594.6	1027.4	2622.0	3.6602	1.6756	5.3358
345	15,541	0.001685	0.009772	1605.5	837.7	2443.2	1631.7	963.4	2595.1	3.7179	1.5585	5.2765
350	16,529	0.001741	0.008806	1642.4	775.9	2418.3	1671.2	892.7	2563.9	3.7788	1.4326	5.2114
355	17,570	0.001808	0.007872	1682.2	706.4	2388.6	1714.0	812.9	2526.9	3.8442	1.2942	5.1384
360	18,666	0.001895	0.006950	1726.2	625.7	2351.9	1761.5	720.1	2481.6	3.9165	1.1373	5.0537
365	19,822	0.002015	0.006009	1777.2	526.4	2303.6	1817.2	605.5	2422.7	4.0004	0.9489	4.9493
370	21,044	0.002217	0.004953	1844.5	385.6	2230.1	1891.2	443.1	2334.3	4.1119	0.6890	4.8009
373.95	22,064	0.003106	0.003106	2015.7	0	2015.7	2084.3	0	2084.3	4.4070	0	4.4070

**Source of Data:** Tables A-4 through A-8 are generated using the Engineering Equation Solver (EES) software developed by S. A. Klein and F. L. Alvarado. The routine used in calculations is the highly accurate Steam\_IAPWS, which incorporates the 1995 Formulation for the Thermodynamic Properties of Ordinary Water Substance for General and Scientific Use, issued by The International Association for the Properties of Water and Steam (IAPWS). This formulation replaces the 1984 formulation of Haar, Gallagher, and Kell (NBS/NRC Steam Tables, Hemisphere Publishing Co., 1984), which is also available in EES as the routine STEAM. The new formulation is based on the correlations of Saul and Wagner (J. Phys. Chem. Ref. Data, 16, 893, 1987) with modifications to adjust to the International Temperature Scale of 1990. The modifications are described by Wagner and Pruss (J. Phys. Chem. Ref. Data, 22, 783, 1993). The properties of ice are based on Hyland and Wexler, "Formulations for the Thermodynamic Properties of the Saturated Phases of H<sub>2</sub>O from 173.15 K to 473.15 K," ASHRAE Trans., Part 2A, Paper 2793, 1983.

**TABLE A-5**

Saturated water—Pressure table

Press., P kPa	Sat. <i>T</i> <sub>sat</sub> °C	Specific volume, m <sup>3</sup> /kg			Internal energy, kJ/kg			Enthalpy, kJ/kg			Entropy, kJ/kg·K		
		Sat. liquid, <i>v</i> <sub>f</sub>	Sat. vapor, <i>v</i> <sub>g</sub>	Sat. liquid, <i>v</i> <sub>f</sub>	Evap., <i>v</i> <sub>f,g</sub>	Sat. vapor, <i>v</i> <sub>g</sub>	Sat. liquid, <i>h</i> <sub>f</sub>	Evap., <i>h</i> <sub>f,g</sub>	Sat. vapor, <i>h</i> <sub>g</sub>	Sat. liquid, <i>s</i> <sub>f</sub>	Evap., <i>s</i> <sub>f,g</sub>	Sat. vapor, <i>s</i> <sub>g</sub>	
1.0	6.97	0.001000	129.19	29.302	2355.2	2384.5	29.303	2484.4	2513.7	0.1059	8.8690	8.9749	
1.5	13.02	0.001001	87.964	54.686	2338.1	2392.8	54.688	2470.1	2524.7	0.1956	8.6314	8.8270	
2.0	17.50	0.001001	66.990	73.431	2325.5	2398.9	73.433	2459.5	2532.9	0.2606	8.4621	8.7227	
2.5	21.08	0.001002	54.242	88.422	2315.4	2403.8	88.424	2451.0	2539.4	0.3118	8.3302	8.6421	
3.0	24.08	0.001003	45.654	100.98	2306.9	2407.9	100.98	2443.9	2544.8	0.3543	8.2222	8.5765	
4.0	28.96	0.001004	34.791	121.39	2293.1	2414.5	121.39	2432.3	2553.7	0.4224	8.0510	8.4734	
5.0	32.87	0.001005	28.185	137.75	2282.1	2419.8	137.75	2423.0	2560.7	0.4762	7.9176	8.3938	
7.5	40.29	0.001008	19.233	168.74	2261.1	2429.8	168.75	2405.3	2574.0	0.5763	7.6738	8.2501	
10	45.81	0.001010	14.670	191.79	2245.4	2437.2	191.81	2392.1	2583.9	0.6492	7.4996	8.1488	
15	53.97	0.001014	10.020	225.93	2222.1	2448.0	225.94	2372.3	2598.3	0.7549	7.2522	8.0071	
20	60.06	0.001017	7.6481	251.40	2204.6	2456.0	251.42	2357.5	2608.9	0.8320	7.0752	7.9073	
25	64.96	0.001020	6.2034	271.93	2190.4	2462.4	271.96	2345.5	2617.5	0.8932	6.9370	7.8302	
30	69.09	0.001022	5.2287	289.24	2178.5	2467.7	289.27	2335.3	2624.6	0.9441	6.8234	7.7675	
40	75.86	0.001026	3.9933	317.58	2158.8	2476.3	317.62	2318.4	2636.1	1.0261	6.6430	7.6691	
50	81.32	0.001030	3.2403	340.49	2142.7	2483.2	340.54	2304.7	2645.2	1.0912	6.5019	7.5931	
75	91.76	0.001037	2.2172	384.36	2111.8	2496.1	384.44	2278.0	2662.4	1.2132	6.2426	7.4558	
100	99.61	0.001043	1.6941	417.40	2088.2	2505.6	417.51	2257.5	2675.0	1.3028	6.0562	7.3589	
101.325	99.97	0.001043	1.6734	418.95	2087.0	2506.0	419.06	2256.5	2675.6	1.3069	6.0476	7.3545	
125	105.97	0.001048	1.3750	444.23	2068.8	2513.0	444.36	2240.6	2684.9	1.3741	5.9100	7.2841	
150	111.35	0.001053	1.1594	466.97	2052.3	2519.2	467.13	2226.0	2693.1	1.4337	5.7894	7.2231	
175	116.04	0.001057	1.0037	486.82	2037.7	2524.5	487.01	2213.1	2700.2	1.4850	5.6865	7.1716	
200	120.21	0.001061	0.88578	504.50	2024.6	2529.1	504.71	2201.6	2706.3	1.5302	5.5968	7.1270	
225	123.97	0.001064	0.79329	520.47	2012.7	2533.2	520.71	2191.0	2711.7	1.5706	5.5171	7.0877	
250	127.41	0.001067	0.71873	535.08	2001.8	2536.8	535.35	2181.2	2716.5	1.6072	5.4453	7.0525	
275	130.58	0.001070	0.65732	548.57	1991.6	2540.1	548.86	2172.0	2720.9	1.6408	5.3800	7.0207	
300	133.52	0.001073	0.60582	561.11	1982.1	2543.2	561.43	2163.5	2724.9	1.6717	5.3200	6.9917	
325	136.27	0.001076	0.56199	572.84	1973.1	2545.9	573.19	2155.4	2728.6	1.7005	5.2645	6.9650	
350	138.86	0.001079	0.52422	583.89	1964.6	2548.5	584.26	2147.7	2732.0	1.7274	5.2128	6.9402	
375	141.30	0.001081	0.49133	594.32	1956.6	2550.9	594.73	2140.4	2735.1	1.7526	5.1645	6.9171	
400	143.61	0.001084	0.46242	604.22	1948.9	2553.1	604.66	2133.4	2738.1	1.7765	5.1191	6.8955	
450	147.90	0.001088	0.41392	622.65	1934.5	2557.1	623.14	2120.3	2743.4	1.8205	5.0356	6.8561	
500	151.83	0.001093	0.37483	639.54	1921.2	2560.7	640.09	2108.0	2748.1	1.8604	4.9603	6.8207	
550	155.46	0.001097	0.34261	655.16	1908.8	2563.9	655.77	2096.6	2752.4	1.8970	4.8916	6.7886	
600	158.83	0.001101	0.31560	669.72	1897.1	2566.8	670.38	2085.8	2756.2	1.9308	4.8285	6.7593	
650	161.98	0.001104	0.29260	683.37	1886.1	2569.4	684.08	2075.5	2759.6	1.9623	4.7699	6.7322	
700	164.95	0.001108	0.27278	696.23	1875.6	2571.8	697.00	2065.8	2762.8	1.9918	4.7153	6.7071	
750	167.75	0.001111	0.25552	708.40	1865.6	2574.0	709.24	2056.4	2765.7	2.0195	4.6642	6.6837	

TABLE A-5

Saturated water—Pressure table (*Concluded*)

Press., P kPa	Specific volume, m <sup>3</sup> /kg			Internal energy, kJ/kg			Enthalpy, kJ/kg			Entropy, kJ/kg·K		
	Sat. temp., <i>T</i> <sub>sat</sub> °C	Sat. liquid, <i>v<sub>f</sub></i>	Sat. vapor, <i>v<sub>g</sub></i>	Sat. liquid, <i>u<sub>f</sub></i>	Evap., <i>u<sub>fg</sub></i>	Sat. vapor, <i>u<sub>g</sub></i>	Sat. liquid, <i>h<sub>f</sub></i>	Evap., <i>h<sub>fg</sub></i>	Sat. vapor, <i>h<sub>g</sub></i>	Sat. liquid, <i>s<sub>f</sub></i>	Evap., <i>s<sub>fg</sub></i>	Sat. vapor, <i>s<sub>g</sub></i>
800	170.41	0.001115	0.24035	719.97	1856.1	2576.0	720.87	2047.5	2768.3	2.0457	4.6160	6.6616
850	172.94	0.001118	0.22690	731.00	1846.9	2577.9	731.95	2038.8	2770.8	2.0705	4.5705	6.6409
900	175.35	0.001121	0.21489	741.55	1838.1	2579.6	742.56	2030.5	2773.0	2.0941	4.5273	6.6213
950	177.66	0.001124	0.20411	751.67	1829.6	2581.3	752.74	2022.4	2775.2	2.1166	4.4862	6.6027
1000	179.88	0.001127	0.19436	761.39	1821.4	2582.8	762.51	2014.6	2777.1	2.1381	4.4470	6.5850
1100	184.06	0.001133	0.17745	779.78	1805.7	2585.5	781.03	1999.6	2780.7	2.1785	4.3735	6.5520
1200	187.96	0.001138	0.16326	796.96	1790.9	2587.8	798.33	1985.4	2783.8	2.2159	4.3058	6.5217
1300	191.60	0.001144	0.15119	813.10	1776.8	2589.9	814.59	1971.9	2786.5	2.2508	4.2428	6.4936
1400	195.04	0.001149	0.14078	828.35	1763.4	2591.8	829.96	1958.9	2788.9	2.2835	4.1840	6.4675
1500	198.29	0.001154	0.13171	842.82	1750.6	2593.4	844.55	1946.4	2791.0	2.3143	4.1287	6.4430
1750	205.72	0.001166	0.11344	876.12	1720.6	2596.7	878.16	1917.1	2795.2	2.3844	4.0033	6.3877
2000	212.38	0.001177	0.099587	906.12	1693.0	2599.1	908.47	1889.8	2798.3	2.4467	3.8923	6.3390
2250	218.41	0.001187	0.088717	933.54	1667.3	2600.9	936.21	1864.3	2800.5	2.5029	3.7926	6.2954
2500	223.95	0.001197	0.079952	958.87	1643.2	2602.1	961.87	1840.1	2801.9	2.5542	3.7016	6.2558
3000	233.85	0.001217	0.066667	1004.6	1598.5	2603.2	1008.3	1794.9	2803.2	2.6454	3.5402	6.1856
3500	242.56	0.001235	0.057061	1045.4	1557.6	2603.0	1049.7	1753.0	2802.7	2.7253	3.3991	6.1244
4000	250.35	0.001252	0.049779	1082.4	1519.3	2601.7	1087.4	1713.5	2800.8	2.7966	3.2731	6.0696
5000	263.94	0.001286	0.039448	1148.1	1448.9	2597.0	1154.5	1639.7	2794.2	2.9207	3.0530	5.9737
6000	275.59	0.001319	0.032449	1205.8	1384.1	2589.9	1213.8	1570.9	2784.6	3.0275	2.8627	5.8902
7000	285.83	0.001352	0.027378	1258.0	1323.0	2581.0	1267.5	1505.2	2772.6	3.1220	2.6927	5.8148
8000	295.01	0.001384	0.023525	1306.0	1264.5	2570.5	1317.1	1441.6	2758.7	3.2077	2.5373	5.7450
9000	303.35	0.001418	0.020489	1350.9	1207.6	2558.5	1363.7	1379.3	2742.9	3.2866	2.3925	5.6791
10,000	311.00	0.001452	0.018028	1393.3	1151.8	2545.2	1407.8	1317.6	2725.5	3.3603	2.2556	5.6159
11,000	318.08	0.001488	0.015988	1433.9	1096.6	2530.4	1450.2	1256.1	2706.3	3.4299	2.1245	5.5544
12,000	324.68	0.001526	0.014264	1473.0	1041.3	2514.3	1491.3	1194.1	2685.4	3.4964	1.9975	5.4939
13,000	330.85	0.001566	0.012781	1511.0	985.5	2496.6	1531.4	1131.3	2662.7	3.5606	1.8730	5.4336
14,000	336.67	0.001610	0.011487	1548.4	928.7	2477.1	1571.0	1067.0	2637.9	3.6232	1.7497	5.3728
15,000	342.16	0.001657	0.010341	1585.5	870.3	2455.7	1610.3	1000.5	2610.8	3.6848	1.6261	5.3108
16,000	347.36	0.001710	0.009312	1622.6	809.4	2432.0	1649.9	931.1	2581.0	3.7461	1.5005	5.2466
17,000	352.29	0.001770	0.008374	1660.2	745.1	2405.4	1690.3	857.4	2547.7	3.8082	1.3709	5.1791
18,000	356.99	0.001840	0.007504	1699.1	675.9	2375.0	1732.2	777.8	2510.0	3.8720	1.2343	5.1064
19,000	361.47	0.001926	0.006677	1740.3	598.9	2339.2	1776.8	689.2	2466.0	3.9396	1.0860	5.0256
20,000	365.75	0.002038	0.005862	1785.8	509.0	2294.8	1826.6	585.5	2412.1	4.0146	0.9164	4.9310
21,000	369.83	0.002207	0.004994	1841.6	391.9	2233.5	1888.0	450.4	2338.4	4.1071	0.7005	4.8076
22,000	373.71	0.002703	0.003644	1951.7	140.8	2092.4	2011.1	161.5	2172.6	4.2942	0.2496	4.5439
22,064	373.95	0.003106	0.003106	2015.7	0	2015.7	2084.3	0	2084.3	4.4070	0	4.4070











**KOLEJ YAYASAN PELAJARAN JOHOR  
SEKOLAH KEJURUTERAAN DAN SAINS KREATIF  
DKM3203 - TERMODINAMIK**

**(1) Hukum Pertama Termodinamik *First Law of Thermodynamics***

$$\sum Q = \sum W \quad Q - W = U_2 - U_1$$

**(2) Proses Alir *Flow Process***

$$\dot{m} = \rho v A = \rho V (\text{kg/s}) = \dot{m} = \frac{CA}{V} \quad Q - W = \dot{m} \left[ (h_2 - h_1) + \left( \frac{C_2^2 - C_1^2}{2} \right) + (Z_2 - Z_1)g \right]$$

$$C = \sqrt{\left\{ 2 \frac{n}{n-1} P_1 V_1 \left( 1 - \left[ \frac{P_2}{P_1} \right]^{(n-1)/n} \right)} \quad Q_f - W_f = \dot{m} \left[ (h_2 - h_1) + \left( \frac{C_2^2 - C_1^2}{2} \right) + (Z_2 - Z_1)g \right]$$

$$V_t = V_1 \left( \frac{P_1}{P_t} \right)^{1/n} \quad A_t = \frac{\bullet m V_t}{C} \quad T_2 = T_1 \left( \frac{P_2}{P_1} \right)^{(n-1)/n}$$

**(3) Sifat Bahan Tulen *Properties of Pure Substance***

**Stim Steam**

$$v = xv_g \quad u = h - Pv \quad h = h_f + xh_{fg} \quad u = u_f + x(u_g - u_f) \quad s = s_f + xs_{fg}$$

**Gas Unggul *Ideal Gas***

$$PV = mRT \quad R = \frac{R_o}{M} \quad R = c_p - c_v \quad \gamma = \frac{c_p}{c_v} \quad c_p = 1.005 \text{ kJ/kg.K}$$

**(4) Proses Tak Alir *Non-Flow Process***

**Proses Isoterma *Isothermal Process* ( $PV = C$ )**

$$U_2 - U_1 = 0 \quad Q = W \quad W = P_1 V_1 \ln \left( \frac{V_2}{V_1} \right) @ W = P_1 V_1 \ln \left( \frac{P_1}{P_2} \right)$$

**Proses Adiabatik *Adiabatic Process* ( $PV\gamma = C$ )**

$$U_2 - U_1 = mc_v(T_2 - T_1) \quad Q = 0 \quad W = \frac{P_1 V_1 - P_2 V_2}{\gamma - 1} = \frac{mR(T_2 - T_1)}{\gamma - 1}$$

$$\text{nisbah kerja} = \frac{\ln \frac{V_2}{V_1} (T_1 - T_2)}{T_1 \ln \frac{V_2}{V_1} + \frac{T_1 - T_2}{\gamma - 1}} \quad \frac{T_2}{T_1} = \left( \frac{P_2}{P_1} \right)^{(n-1)/n} = \left( \frac{V_1}{V_2} \right)^{\gamma-1} \quad \eta_{\text{ad, rev}} = 1 - \frac{T_1}{T_2}$$

**(5) Hukum Kedua Termodinamik *Second Law of Thermodynamics***

**Enjin Haba *Heat Engine***

$$\eta_{\text{ih}} = \frac{W_{\text{net, net}}}{Q_h} = 1 - \frac{Q_L}{Q_h}$$

**Pam Haba *Heat Pump***

$$COP_{\text{ih, rev}} = \frac{T_h}{T_h - T_L} = \frac{1}{1 - \frac{T_L}{T_h}}$$